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Synthesis and Characterisation of a Paramagnetic [2]Rotaxane Based on a Crown Ether-Like Wheel  
Incorporating a Nitroxide Motif

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# Synthesis and characterization of a paramagnetic [2]rotaxane based on a crown ether-like wheel incorporating a nitroxide motif

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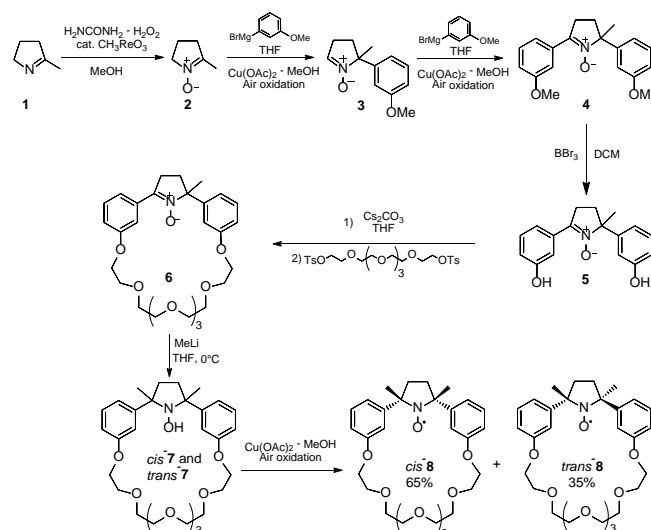
**Abstract:** The synthesis of a new nitroxide crown ether (**8**) and its use as the wheel in a bistable [2]rotaxane containing dialkylammonium and 4,4'-bipyridinium recognition sites, is reported. The synthesis of **8** was achieved by the sequential addition of substituted phenyl groups to nitron derivatives leading to the preferential formation of the *cis*-stereoisomer. Due to charge-dipolar interactions between the nitroxide unit and bipyridinium moiety, it was possible to probe the movement of the macrocycle between the two molecular stations of the [2]rotaxane after addition of a base by measuring the nitrogen hyperfine splittings in the corresponding EPR spectra. The equilibrium constant for the complexation of dibenzyl viologen by macrocycle **8** was also determined by EPR titration.

Nitroxide radicals are an important group of open-shell molecules with intriguing catalytic and magnetic properties, giving rise to diverse and important applications in many fields of chemistry and related sciences. These applications include the biomedical field (diagnostic spin probes,<sup>[1]</sup> spin traps,<sup>[2]</sup> and potential antioxidant therapeutics),<sup>[3]</sup> the development of sustainable chemical processes (catalysts in aerobic oxidation)<sup>[4]</sup> and material chemistry (radical initiators for living and other controlled polymerizations).<sup>[5]</sup>

However, with the notable exception of the alkyl-nitronyl nitroxide organic magnet systems,<sup>[6]</sup> the use of nitroxide in molecular-level assembly processes remains largely unexplored, certainly in comparison with the use of ethers, amides, imines, ureas, and guanidinium units.<sup>[7]</sup> This is somewhat surprising given their desirable intrinsic properties: nitroxides possess one of the largest dipole moments known for any functional-group (close to 3 D),<sup>[8]</sup> making them potentially useful for control of molecular orientation: the N-O• motif is a good hydrogen<sup>[9]</sup> and halogen-bond<sup>[10]</sup> acceptor and they offer a simple reversible route to diamagnetic species via one-electron electrochemical oxidation.<sup>[11]</sup> Rotaxanes and catenanes containing nitroxide units have been described in the literature,<sup>[12]</sup> but in all cases the paramagnetic unit was used as spin probe and linked to the macrocycle or to the thread without participating to the recognition process. In most of these examples spin exchange

interaction between two radical units was used to probe the shuttling process of the macrocycle between two different recognition sites triggered by protonation or electron transfer.<sup>[12]</sup> While the dimerization of radical ions has been proposed to move a macrocyclic ring between different molecular recognition sites in a [2]rotaxane,<sup>[13]</sup> to the best of our knowledge no examples of mechanical interlocked molecules like rotaxanes or catenanes in which a nitroxide functionality could be used as recognition unit has been reported so far.

Based on this, we decided to prepare and characterize the first example of a nitroxide-containing rotaxane in which the N-O• motif acts not only as EPR probe for the shuttling process, but could be also directly involved in the recognition process of a molecular site. This was achieved by inserting a dialkyl nitroxide functionality in a crown ether-like frame containing seven ethereal oxygen atoms (**8**, Scheme 1) and by using this macrocycle for complexation of cationic guests. We will show that the EPR read-out discriminates between dialkylammonium (NH<sub>2</sub><sup>+</sup>) and 4,4'-bipyridinium (BPY<sup>2+</sup>) recognition sites and this favourable feature can be used to probe shuttling process of the paramagnetic wheel in a bistable [2]rotaxane.



**Scheme 1.** Reaction scheme for the preparation of the paramagnetic macrocycle **8**.

Nitroxide based crown ether-like macrocycles are not new and were reported as far back as 1983 by Keana and co-workers.<sup>[14]</sup> In such examples, the crown ether is inserted in both the aromatic rings at the *ortho* position in a 2,5-dimethyl-2,5-diphenylpyrrolidine-N-oxyl radical. In order to reduce the steric hindrance around the nitroxidic functionality, we decided to

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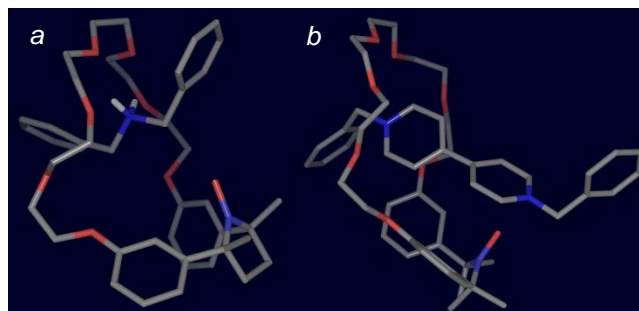
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prepare a new derivative (**8**) where the two aromatic rings are linked *via* a *meta*-substitution to the nitroxide and to the polyether moieties. By adapting the original procedure proposed by Keana *et al.*<sup>[14]</sup> for the preparation of the *ortho*-derivative, we prepared the new *meta*-substituted macrocycle **8** as described in Scheme 1. An important point which deserves a comment concerns the reaction of nitrene **6** with MeLi that provided hydroxylamine derivative **7** in quantitative yield as a mixture of *cis* and *trans* isomers in 6.5:3.5 molar ratio, as assessed by <sup>1</sup>H NMR (see Supporting Information). Although we were not able to obtain a single crystal of **7** for X-ray structure determination, we attributed the stereochemistry of the main isomer to the *cis* form by following the model presented by Shibata *et al.* reporting that an organometallic reagent attacks the pyrrolidine nitrene from the less hindered side.<sup>[15]</sup> In our case, MeLi is expected to add to nitrene **6** from the side opposite to the bulky phenyl group to give the *cis*-hydroxylamine **7**.

The mixture of the two isomers (*cis*- and *trans*-**7**) was oxidized with air in the presence of Cu<sup>2+</sup> ions. After purification of the crude mixture by column chromatography, a main fraction containing the nitroxide isomer *cis*-**8** was recovered together with less abundant fractions corresponding to the other isomer of nitroxide (*trans*-**8**) and unreacted *trans*-**7**. Assignment of the *cis*-stereochemistry to the main fraction was obtained by NMR analysis after its reduction with phenylhydrazine.<sup>[16]</sup> Interestingly, the two nitroxide isomers show different relative rates of reduction. While *cis*-**8** isomer can be reduced to the corresponding hydroxylamine in 10-15 minutes, the *trans*-isomer is much more resistant toward reduction, requiring several days. This observation is in agreement with a previous report on *trans*-azethoxyl acid which was found to be the most resistant toward reduction if compared to analogous *cis*-derivatives.<sup>[17]</sup> These findings corroborate the hypothesis previously exposed on the attribution of the stereoisomers. On this basis, the presence of unreacted *trans*-**7** after oxidation of the mixture of hydroxylamines **7** suggests that also this step occurs more efficiently for the *cis* than the *trans* isomer because of the less steric hindrance around the N-OH functionality in the former one. The CH<sub>2</sub>Cl<sub>2</sub> solution of macrocycle *cis*-**8** shows a typical nitroxide three line EPR spectrum and is characterized by  $a_N = 14.35$  G and  $g = 2.0061$ . EPR analysis of the CH<sub>2</sub>Cl<sub>2</sub> solution containing *trans*-**8** was virtually identical, apart a very small difference in the nitrogen coupling ( $a_N = 14.39$  G, see Supporting Information).

The host properties of the most abundant isomer, *cis*-**8**, towards cations which are known to be selectively complexed by crown ether derivatives, were then investigated by EPR. Na<sup>+</sup> and K<sup>+</sup> picrate salts are insoluble in CH<sub>2</sub>Cl<sub>2</sub>. However when a solution containing an excess of *cis*-**8** was added to the suspension complete solubilisation of metal salts was observed. This clearly suggests that the paramagnetic macrocycle *cis*-**8** maintains the complexation ability of crown ethers towards alkaline cations. Unexpectedly, the EPR spectra of the nitroxide *cis*-**8** recorded in the presence of alkaline metal cations did not show any significant variation in the spectroscopic parameters compared to those measured in the absence of charged guests. The striking conclusion is that despite the position of the nitroxide group with respect to the ion binding cavity, the radical unit is not participating to the complexation process of the metal

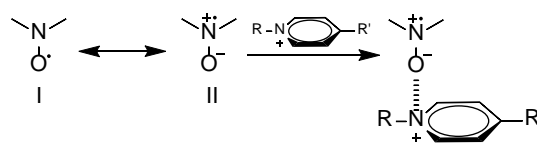
cation (which thus is expected to interact only with crown ether moiety, maintaining the cation far from the nitroxide). A similar result was obtained in the presence of dibenzyl ammonium PF<sub>6</sub> salt which did not produce a significant change in the EPR spectrum of *cis*-**8**. Stochastic dynamic (SD) simulations performed *in vacuo* at 298 K on the dibenzylammonium@*cis*-**8** complexes by using the AMBER\* force field (see Supporting Information), suggested that the radical unit is indeed not participating to the complexation process of the ammonium group (see Figure 1a).



**Figure 1.** Stochastic dynamic simulations snapshots providing the possible structures of dibenzylammonium@*cis*-**8** (a) and DBV@*cis*-**8** (b) complexes.

A different spectroscopic behaviour was instead observed when the EPR spectrum of *cis*-**8** was recorded in the presence of bis-cation dibenzyl viologen (DBV) as the PF<sub>6</sub> salt. In this case the complete solubilization of the bis-pyridinium salt was accompanied by an increase of the nitrogen coupling ( $a_N$ ) value in the EPR spectrum of *cis*-**8**. As an example, in acetone at room temperature  $a_N$  increases from 14.17 G to 14.32 G in the presence of 0.042 M DBV.

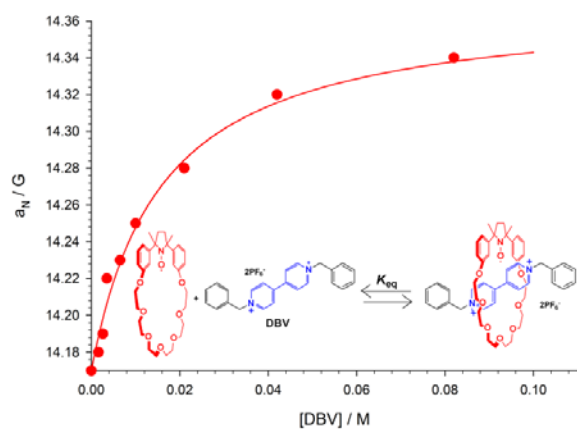
Although the nature of the interaction is not fully defined, SD calculations (Figure 1b) suggests that binding of the bis-pyridinium salt occurs such that in the complex, one of the two cationic portions is interacting with the nitroxide unit by a charge-dipolar interaction. In this model, the increase of  $a_N$  can be explained on the basis of the two canonical forms I and II of the nitroxide function (Scheme 2). The greater the interaction with the positive charge of pyridinium unit, the more favoured will be structure II, bearing larger electron density on the oxygen and spin density on the nitrogen atom.



**Scheme 2.** Charge-dipole interaction between a dialkyl nitroxide and pyridinium unit.

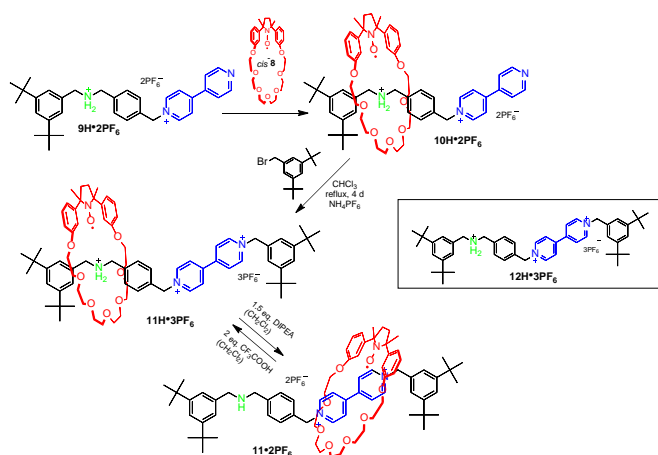
As shown in Figure 2 the changes in the value of  $a_N$  as a function of DBV concentration were well-modelled in acetone at 298 K by a 1:1 binding model using standard curve fitting methods and  $K_{eq}$  was found to be  $69 \pm 19$  M<sup>-1</sup>.<sup>[18]</sup> This value is

very close to that ( $K_{eq} = 82 \text{ M}^{-1}$ ) measured in ACN at 298 K for the complexation of DBV by the closely related diamagnetic macrocycle dibenzo[24]crown-8.<sup>[19]</sup>



**Figure 2.** Plot of the EPR nitrogen coupling,  $a_N$ , in *cis-8* (0.05 mM) versus the concentration of DBV at 298 K in acetone. The line represents the theoretical dependence of  $a_N$  on DBV concentration calculated by taking into account the formation of 1:1 complex and  $K_{eq}=69 \text{ M}^{-1}$ .

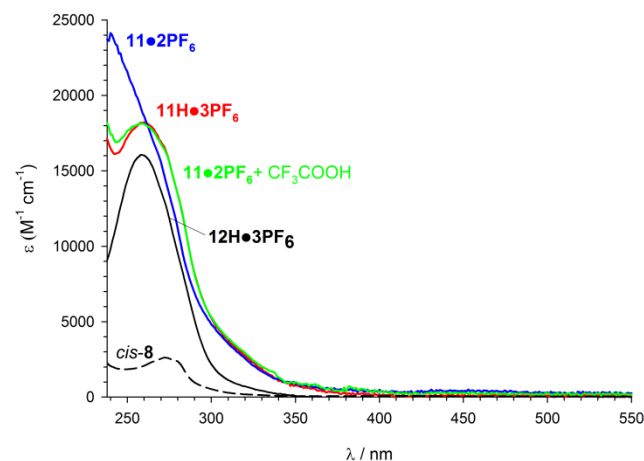
Because the EPR read-out discriminates between different cationic guests we evaluated the possibility of using radical macrocycle *cis-8* in a bistable [2]rotaxane containing dialkylammonium and 4,4'-bipyridinium recognition sites in the thread. The rotaxane **11H•3PF<sub>6</sub>** (see Scheme 3) was prepared by following the classical threading-stopping approach, which involves the in situ self-assembly of a pseudorotaxane, **10H•2PF<sub>6</sub>**, between the ring *cis-8* and the half-dumbbell intermediate **9H•2PF<sub>6</sub>**.



**Scheme 3.** Reaction scheme for the preparation of the rotaxane **11H•3PF<sub>6</sub>** and the switching process between **11H•3PF<sub>6</sub>** and **11•2PF<sub>6</sub>**.

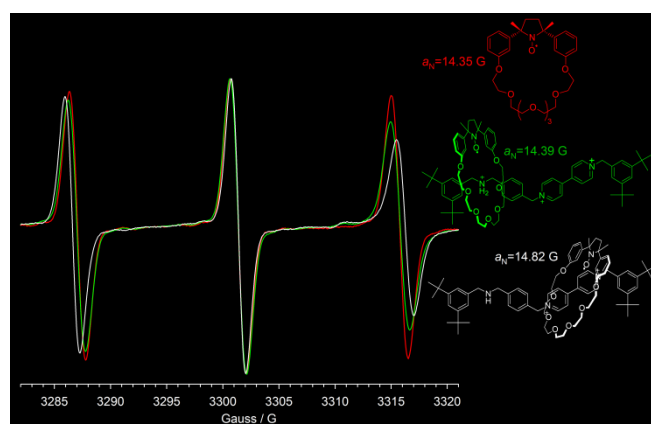
Evidence for the formation of pseudorotaxane in  $\text{CHCl}_3$  was obtained by observing complete solubilisation of the half-

dumbbell **9H•2PF<sub>6</sub>** which is almost completely insoluble in chlorinated hydrocarbons. The pseudorotaxane was then reacted with (3,5-di-*tert*-butyl)-benzyl bromide by  $\text{S}_{\text{N}}2$ -mediated stopping reaction to afford rotaxane **11H•3PF<sub>6</sub>** in 30% yield. The rotaxane formation was detected by ESI-MS (see Supporting Information), UV-visible and EPR spectroscopy.



**Figure 3.** Absorption spectra of macrocycle *cis-8* (dashed black line), free dumbbell **12•2PF<sub>6</sub>** (black line) and rotaxanes **11H•3PF<sub>6</sub>** before (red line) and after subsequent addition of 1.5 eq of base (blue line) and  $\text{CF}_3\text{COOH}$  (green line) in ACN at room T.

Figure 3 compares the absorption spectra of **11H•3PF<sub>6</sub>** and its constituents, i.e., the free dumbbell **12H•3PF<sub>6</sub>** and the macrocycle *cis-8*. The most intense absorption feature in the rotaxane arises from the viologen unit ( $\epsilon = 18230 \text{ M}^{-1} \text{ s}^{-1}$ ,  $\lambda = 259 \text{ nm}$ ). This band overlaps with the weaker band due to the aromatic rings of the macrocycle unit and it is essentially unperturbed on going from **12H•3PF<sub>6</sub>** to **11H•3PF<sub>6</sub>**. This spectral feature suggests that the macrocycle is not significantly interacting with the  $\text{BPY}^{2+}$  unit, consistent with it being localised predominantly over the ammonium site of the thread as expected.



**Figure 4.** EPR spectra of macrocycle *cis-8* (red line) and rotaxanes **11H•3PF<sub>6</sub>** (green line) and **11•2PF<sub>6</sub>** (white line) in  $\text{CH}_2\text{Cl}_2$  at room T.

Figure 4 (green line) shows the EPR spectrum of rotaxane **11H•3PF<sub>6</sub>**, recorded in CH<sub>2</sub>Cl<sub>2</sub> at 298 K, exhibiting typical nitroxide EPR signals with the high field line broadened due to restricted tumbling and, thus, incomplete averaging of the anisotropic components of the hyperfine and *g*-tensor. The *a<sub>N</sub>* hyperfine splitting and the rotational correlation times  $\tau$ , are reported in Table 1.<sup>[20]</sup> The value of *a<sub>N</sub>* = 14.39 G is very close to that observed both for the free macrocycle and the dibenzylammonium@*cis*-**8** complex. According to UV analysis, these EPR data strengthens the conclusion that the macrocycle predominantly encircles the ammonium station. The higher affinity shown by *cis*-**8** for the complexation of dibenzyl ammonium site compared to the BPY<sup>2+</sup> one, is in line with the complexation behaviour shown by the related macrocycle dibenzo[24]crown-8 dibenzyl for which the association constant with dibenzyl ammonium cation is about 5 times higher (*K<sub>eq</sub>* = 420 M<sup>-1</sup>)<sup>[21]</sup> than that measured with dibenzyl viologen (*K<sub>eq</sub>* = 82 M<sup>-1</sup>)<sup>[19]</sup> in ACN at 298K.

The rotational correlation time measured in the EPR spectrum of **11H•3PF<sub>6</sub>** is 4.3 times higher than that measured in the spectrum of *cis*-**8** (Table 1). This substantial reduction of the tumbling rate is typical of spin labels attached to larger molecules and can be considered a further indication of the formation of the MIM structure.

**Table 1.** EPR parameters of nitroxides in CH<sub>2</sub>Cl<sub>2</sub> at 298K.<sup>[20]</sup>

Nitroxide	<i>a<sub>N</sub></i> / G	Line width of central line/ G	Signal ratio <i>h<sub>0</sub>/h<sub>-1</sub></i>	Correlation time $\tau$ / s
<i>cis</i> - <b>8</b>	14.35	1.38	1.10	4.6×10 <sup>-11</sup>
<b>11H•3PF<sub>6</sub></b>	14.39	1.42	1.47	2.0×10 <sup>-10</sup>
<b>11•2PF<sub>6</sub></b>	14.82	1.41	1.65	2.6×10 <sup>-10</sup>

Although, we found some differences comparing the <sup>1</sup>H-NMR data of the dumbbell and the rotaxane,<sup>[22]</sup> NMR studies are not particularly useful for the characterization of **11H•3PF<sub>6</sub>** because most of the signals were broadened by the presence of the paramagnetic center. It is worth noting that attempts to reduce the rotaxane with phenylhydrazine to the diamagnetic hydroxylamine derivative were unsuccessful, denoting a considerable inertness of the nitroxidic interlocked assembly to this reaction conditions.

Thus, the shuttling of the macrocycle towards the secondary BPY<sup>2+</sup> station (Scheme 3) was followed by UV-Visible and EPR spectroscopies. Treatment of rotaxane **11H•3PF<sub>6</sub>** in CH<sub>2</sub>Cl<sub>2</sub> with the non-nucleophilic diisopropyl-ethylamine (*i*-Pr<sub>2</sub>EtN, DIPEA), which is strong enough to deprotonate the NH<sub>2</sub><sup>+</sup> center,<sup>[23]</sup> led to perturbation of the dominant viologen UV-absorption features around 259 nm. As shown in Figure 3 (blue line), deprotonation is accompanied by a small blue shift and an increase in the absorbance of viologen band. This behavior is consistent with the displacement of *cis*-**8** to predominantly encircle the bipyridinium station leading to a stabilization of the ground state of BPY<sup>2+</sup> due to charge-dipolar interactions between this unit and the nitroxidic moiety. In contrast,

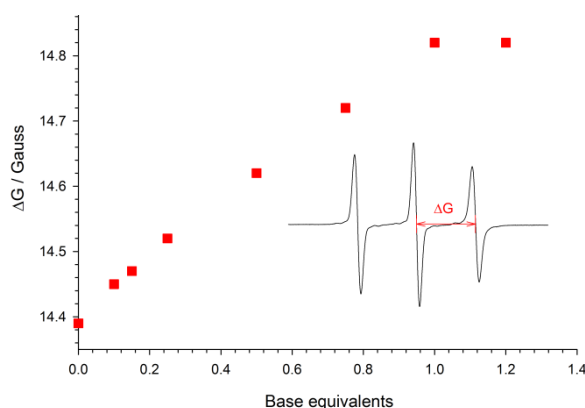
deprotonation of free dumbbell **12H•3PF<sub>6</sub>** did not cause appreciable variation in the corresponding absorption band (see Supporting Information).<sup>[24]</sup>

Finally, the addition of a stoichiometric amount of trifluoroacetic acid (TFA) to the rotaxane after the addition of the base caused the quantitative recovery of the initial absorption spectrum (see Figure 3, green line), consistent with the system returning to its state in which macrocycle predominantly encircles the ammonium station.

Addition of increasing amount of *i*-Pr<sub>2</sub>EtN to the rotaxane also produces a noticeable increase in the EPR nitrogen coupling which reaches a plateau after addition of 1 eq. of base (*a<sub>N</sub>* = 14.82 G,  $\Delta a_N = +0.43$  G, see Figure 4, white line). The increase in the value of *a<sub>N</sub>* is similar to that observed when free DBV is complexed by the macrocycle *cis*-**8** (see Figure 2).

During the titration with *i*-Pr<sub>2</sub>EtN, the protonated (**12H•3PF<sub>6</sub>**) and deprotonated (**12•2PF<sub>6</sub>**) forms of the rotaxane are present in solution simultaneously and in equilibrium with one another. It has been reported that, in a related diamagnetic rotaxane, base induced shuttling of a dibenzo[24]crown ether macrocycle to BPY<sup>2+</sup> station is complete in seconds and acid-induced reverse of this process occurs on a timescale of tens of milliseconds.<sup>[25]</sup> If we assume proton transfer between **12H•3PF<sub>6</sub>** and **12•2PF<sub>6</sub>** takes place on similar time scales the exchange between the protonated and deprotonated states will be significantly slower than the EPR timescale. Thus, during the titration, the EPR spectrum is expected to be the superposition of the two signals for **12H•3PF<sub>6</sub>** and **12•2PF<sub>6</sub>**. However, due to the large spectral line width, the detection of separate EPR signals is prevented and the variation in the relative amount of the radical co-conformations is expected to give rise only to a variation in the corresponding apparent field separation of the spectral lines.

Indeed, Figure 5 shows the changes in the field separation,  $\Delta G$ , between the central and high field lines, observed experimentally by EPR, upon addition of increasing amounts of *i*-Pr<sub>2</sub>EtN to **11H•3PF<sub>6</sub>**. When *i*-Pr<sub>2</sub>EtN is less than 1 eq. (referred to the rotaxane) both nitroxides **11H•3PF<sub>6</sub>** and **11•2PF<sub>6</sub>** are present and  $\Delta G$  reflects the relative amount of protonated and deprotonated species.



**Figure 5.** Field separation,  $\Delta G$ , between the central and high field lines, monitored by EPR, upon addition of *i*-Pr<sub>2</sub>EtN to **11H•3PF<sub>6</sub>**.

The switching of the rotaxane also leads to a slight increase of rotational correlation time  $\tau$  from  $2.0 \times 10^{-10}$  to  $2.6 \times 10^{-10}$  s (see Figure 4 and Table 1). The change in the rotational correlation time when passing from one conformation to the other can be associated to the reduced ability of the 'wheel' to rotate relative to the thread when is located at the BPY<sup>2+</sup> station.

Conversely, the base addition performed on a 1:1 mixture of the non-interlocked rotaxane components (*cis*-**8** and the free dumbbell, **12H•3PF<sub>6</sub>**) did not cause variations in the nitrogen splitting nor in the correlation time. Such a control experiment confirms that the observed EPR spectral parameters changes arise from the interaction between the radical unit and the bipyridinium unit which become possible as a result of ring shuttling in the deprotonated rotaxane.

The addition of a stoichiometric amount of TFA after the addition of *i*-Pr<sub>2</sub>EtN caused the quantitative recovery of the initial EPR spectrum ( $a_N = 14.39$  G, see Supporting Information) as a consequence of the macrocycle returning to predominantly encircle the ammonium station. The complete base- and acid-induced switching cycle of the EPR pattern was repeated several times without an appreciable loss of signal, highlighting the reversibility of the process.

## Conclusion

In conclusion, we have reported the synthesis of the first nitroxide rotaxane in which the nitroxide unit acts not only as EPR probe for the switching, but interacts with a recognition site (bipyridinium) in the thread. While in previous work, the shuttling process was detected by EPR because of switching between non-coupled (three-line EPR spectrum) and coupled (five-line EPR spectrum) states of nitroxide radicals connected to the ring and dumbbell, in this work we have shown that a simple change in the nitrogen coupling reports on switching process between dialkylammonium and 4,4'-bipyridinium recognition sites. The nitroxide behaves as spin probe of shuttling process because is able to interact with bipyridinium unit by charge-dipolar interaction. Finally, the presence of nitroxide units provides an unusual functional addition to the diversity of templates available for assembling rotaxanes and could potentially lead to new generations of radical molecular machines and superstable radicals.

## Acknowledgements

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**Keywords:** rotaxane • nitroxide • electron paramagnetic resonance • spin probes • supramolecular chemistry

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