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# Selective enhancement of organic dye properties through encapsulation in rotaxane orientational isomers†

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We report the synthesis and characterisation of [2]rotaxanes based on a stilbazolium dye and a calix[6]arene macrocycle. Since both components are non-symmetric, two orientational isomers are obtained. The two isomers display distinct photophysical and photochemical properties in solution and solid state, superior to the unencapsulated dye.

Stilbazolium salts are push-pull dyes comprising a pyridinium electron acceptor and an electron donor unit bridged by a double bond. They are studied as fluorescent sensors<sup>1–3</sup> and non-linear optical materials;<sup>4–7</sup> moreover, depending on their substituents, they can undergo *E-Z* photoisomerisation around the central double bond.<sup>8,9</sup> On account of the charge-transfer character of their electronic transitions,<sup>10</sup> the photophysical properties of these cationic dyes strongly depend on their environment, and they can be affected by encapsulation within the cavity of hosts in supramolecular complexes.<sup>11–13</sup> In this regard, recently, some of us have studied the association of a stilbazolium axle within the cavity of a calix[6]arene macrocycle.<sup>14</sup> Since the host and the guest are non-symmetric, the resulting pseudorotaxane complex displays two possible orientational isomers. In solution, these isomers are both present and exhibit different emission spectra. The ratio between the two isomers can be adjusted by modifying the temperature, but their continuous interconversion prevents their separation. In this study, we investigate rotaxanes **1-Down** and

**1-Down**, wherein the stilbazolium-based axle **2** is encapsulated in the calix[6]arene **3** wheel, which directs its upper or lower rim towards the 3,5-dimethyl piperidine unit of the axle, respectively (Fig. 1). Thanks to the mechanical bond, the separation of the isomers is achieved, allowing a thorough investigation of the system in solution and in the solid state. Although different examples of rotaxane orientational isomers are reported in the literature,<sup>15–20</sup> usually they exhibit similar photophysical properties.<sup>15,20</sup> On the other hand, the encapsulation inside the

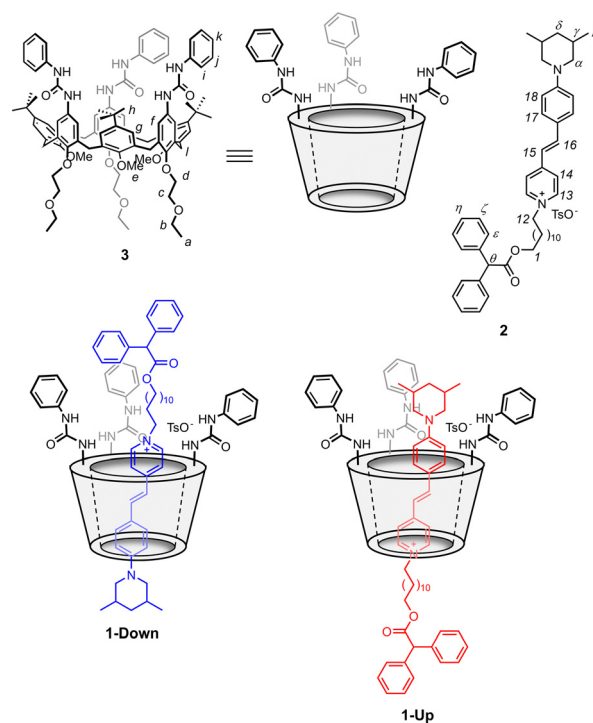


Fig. 1 Chemical structure of axle **2** and the [2]rotaxane orientational isomers **1-Down** and **1-Up** (letters and numbers on the structure of **2** and **3** are used for NMR assignment).

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Table 1 Photophysical and photochemical data of the investigated compounds

Compound	Toluene						Solid		
	$\lambda_{\text{abs}}/\text{nm}$	$\lambda_{\text{em}}/\text{nm}$	$\Phi_{\text{fluo}}$ (%)	$\Phi_{EZ}$ (436 nm) (%)	$\Phi_{ZE}$ (436 nm) (%)	$k_{Z \rightarrow E}^b/\text{s}^{-1}$	$\lambda_{\text{abs}}/\text{nm}$	$\lambda_{\text{em}}/\text{nm}$	$\Phi_{\text{fluo}}$ (%)
<b>2<sup>a</sup></b>	450	618	18	14	42	$3 \times 10^{-4}$	451	613	1
<b>1-Up</b>	468	561	83	7	71	$1 \times 10^{-4c}$	463	587	11
<b>1-Down</b>	425	533	20	36	41	$\sim 10^{-7}$	440	576	32

<sup>a</sup> Properties determined in the presence of an excess of tosylate counterion (see ESI). <sup>b</sup> Rate constant of the thermal  $Z \rightarrow E$  isomerisation at room temperature. <sup>c</sup> A minor secondary process is observed; see ESI for details.

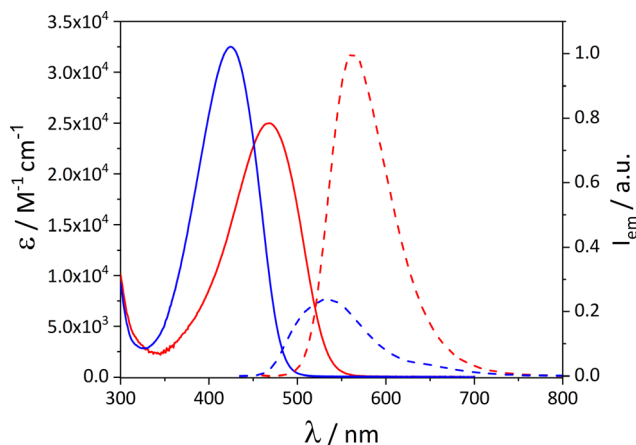


Fig. 3 Absorption (full lines) and emission spectra (dashed lines) of toluene solutions of **1-Up** (red lines,  $\lambda_{\text{ex}} = 454$  nm) and **1-Down** (blue lines,  $\lambda_{\text{ex}} = 425$  nm). The emission spectra intensity is proportional to the fluorescence quantum yield (0.83 for **1-Up** and 0.20 for **1-Down**).

the two isomers show distinct spectra (Fig. 3 and Tables 1), and **1-Up** is characterised by a fluorescence quantum yield of 83%, *i.e.*, four times larger than **1-Down** and ion-paired **2**.

A rationalisation of these effects is not straightforward, nevertheless, a qualitative explanation of the observed phenomena is attempted. As for the absorption properties, stilbazolium ions exhibit negative solvatochromism,<sup>10</sup> *i.e.* a blue-shifted absorption on increasing the polarity of the media, which would suggest that in **1-Down** the calixarene host provides a more polar environment to the guest than in **1-Up**.

With regard to the emission properties, the smaller Stokes shift and larger quantum yield of **1-Up** can possibly be accounted for by a deeper encapsulation of the stilbazolium and/or a more rigid environment. Indeed, the ground and excited states exhibit different polarities,<sup>10</sup> therefore the energy of the emission can be influenced by the solvent reorganisation. A deeper encapsulation would minimise this effect, as the cavity of the macrocycle may screen the solvent. Moreover, stilbazolium fluorescence can benefit from the presence of an environment that limits molecular rotations.<sup>11</sup>

The influence of encapsulation on the properties of the stilbazolium unit was also investigated in the solid state. Indeed, the inclusion of stilbazolium fluorophores in a restricted environment, such as a macrocyclic cavity<sup>6,7</sup> or a molecular organic framework (MOF),<sup>4,5</sup> is known to prevent the intermolecular  $\pi$ - $\pi$  stacking responsible for the aggregation caused quenching (ACQ)<sup>21</sup> and

enhance the emission in the solid state. The absorption and the emission spectra of the solid compounds resemble the ones observed in solution (see Fig. S35 and S36, ESI†) and, in particular, the two orientational isomers show distinct spectra (Table 1). The emission of **2** is weak ( $\Phi_{\text{f}} = 1\%$ ) due to ACQ, while **1-Up** and **1-Down** exhibit a stronger fluorescence ( $\Phi_{\text{f}} = 11\%$  and  $32\%$ , respectively), as a consequence of the encapsulation of the stilbazolium unit inside the calix[6]arene cavity. To the best of our knowledge, this is the first example of such an enhancement obtained through a rotaxane architecture.

The larger quantum yield of **1-Down** with respect to **1-Up** (an opposite outcome with respect to what was observed in solution) is possibly related to the conformation assumed by the rotaxane in the solid state. To further investigate this aspect, we attempted the crystallisation of the rotaxanes. We successfully isolated crystals suitable for X-ray diffraction analysis from the diffusion of hexane into a toluene solution of **1-Down**. The solved X-ray structure confirms the identity of this orientational isomer (Fig. 4 and Fig. S53–S56, ESI†).§ Considerable effort has been devoted to growing crystals of the **1-Up** isomer, but the size and quality necessary for X-ray diffraction experiments could not be achieved, thus preventing further considerations.

The photoisomerisation processes of axle **2** and rotaxanes **1** were investigated through UV-Vis and NMR spectroscopy (see ESI†). Upon irradiation of the ion-paired **2** at 436 nm, a modest decrease of the absorption band, ascribed to the  $E$  to  $Z$  photoisomerisation reaction, is observed (see Fig. S43, ESI†). Upon keeping the solution

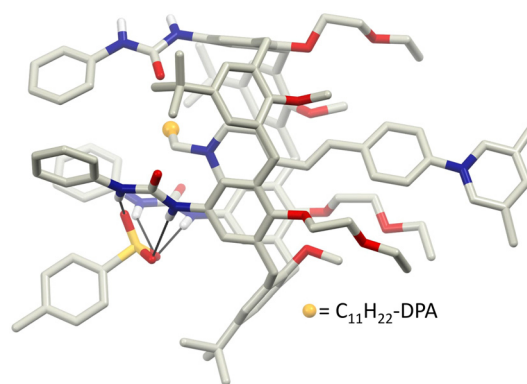


Fig. 4 X-Ray molecular structure of **1-Down** showing the hydrogen bond interactions between the phenylureas N–H groups and the tosylate anion. The disordered alkyl chain bearing the diphenylacetic stopper is depicted as a gold sphere, while hydrogen atoms, except those involved in hydrogen bonding, have been omitted for clarity (see ESI† for further details).



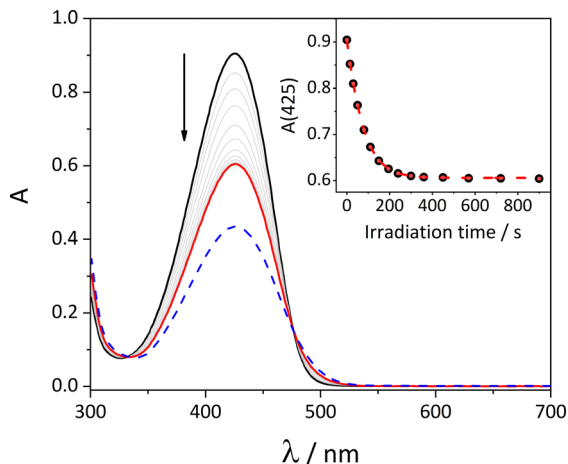


Fig. 5 Absorption spectra of a toluene solution of *E* **1-Down** ( $2.8 \times 10^{-5}$  M, black to red lines) upon irradiation at 436 nm. The calculated spectrum of *Z* **1-Down** is reported for comparison (blue dashed line). Inset: Absorption changes at 425 nm (black dots) together with the fitting of the data (red dashed line).

in the dark, the thermal back isomerisation occurs, leading to the recovery of the initial spectrum within a few hours, with a first-order kinetic constant  $k_{Z \rightarrow E} = 3 \times 10^{-4} \text{ s}^{-1}$  (see Fig. S46, ESI<sup>†</sup>). The photoisomerisation quantum yields were obtained by fitting the photokinetic profile (Table 1). Also the photochemistry of the stilbazolium moiety is influenced by encapsulation within the calixarene cavity, with a dependence on the orientation of the macrocycle. Indeed, the *E* to *Z* photoisomerisation quantum yield  $\Phi_{EZ}$  of **1-Up** is 7%, and its photostationary state (PSS) at 436 nm contains only 10% of *Z* isomer (Table 1). On the other hand, **1-Down** is characterised by a  $\Phi_{EZ}$  of 36% and a PSS at 436 nm that contains 64% of the *Z* isomer (Fig. 5 and Table 1). Moreover, the thermal back isomerisation of **1-Down** is much slower with respect to the free axle in solution and to the other rotaxane isomer, and its value could only be estimated ( $k_{Z \rightarrow E} \sim 10^{-7} \text{ s}^{-1}$ , see Fig. S47, ESI<sup>†</sup>). The photoisomerisation quantum yields are in accordance and complementary with the emission quantum yields. The results obtained by NMR spectroscopy are consistent with the UV/Vis studies (Fig. S49–S52, ESI<sup>†</sup>). It is worth to highlight the signals appearing between 0 and  $-1$  ppm in the  $^1\text{H}$  NMR spectrum of **1-Down** upon photoisomerization: such signals could be assigned to the protons of the alkyl chain shielded by the cavity of the wheel; this might suggest that the isomerization of the stilbazolium unit causes a partial displacement of the macrocycle, forcing the wheel to move toward the opposite end of the axle.

Overall, the encapsulated compound shows an improvement of either the photophysical or photochemical properties with respect to the free axle in solution, with a remarkable and distinctive dependence on the orientation of the macrocycle. Indeed, the “up” orientation favours the radiative decay of the stilbazolium dye, with an emission quantum yield as high as 83%, whereas the “down” orientation fosters the photoisomerisation reaction, leading to the kinetically inert *Z* isomer. This peculiar feature allows to select one property (emission or photoisomerisation) over the other, tailoring the system depending on the desired application.

Remarkably, in this system, the orientation plays a role in the outcome of the properties that is as important as the encapsulation itself. Moreover, the orientational isomers display different emission spectra both in solution and solid state, and this feature gives access to a straightforward way to identify the orientation of the macrocycle. This information is trivial in the present rotaxanes, but it can be crucial when studying more complex mechanically interlocked molecules, in which the application of other techniques (such as X-ray crystallography and NMR spectroscopy) could be limited or their interpretation ambiguous. Indeed, understanding the orientation of the macrocycle can be pivotal to get insights on its movements between different stations, a fundamental aspect in the study of artificial molecular machines.

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## Conflicts of interest

There are no conflicts to declare.

## Notes and references

§ Upon mixing **3** and **4** in  $\text{C}_6\text{D}_6$  at room temperature, the orientational pseudorotaxane isomer **P-Up** is formed preferentially (**P-Up**: **P-Down** = 82: 18). The mixture reaches its equilibrium (48:52) in *ca.* 2 h (see ESI<sup>†</sup>).

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